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Analytica Chimica Acta	134	135/1	135/2	136	137	138	139	140/1 140/2	141	142	143	144

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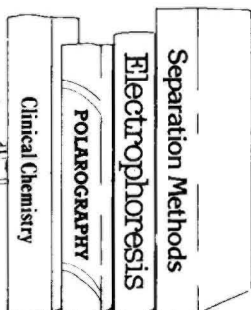
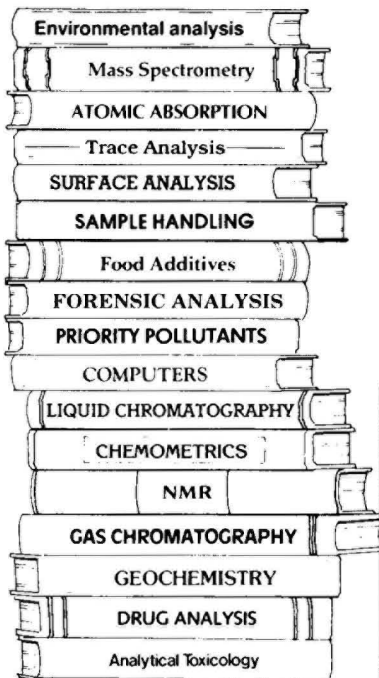
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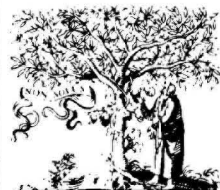
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Hg <sup>2+</sup> added ( $\mu\text{g}$ )	1.0	2.0	3.0	5.0
Extraction (%)	95.0	99.8	99.5	89.0

Experimental information which is relevant to all the results in the table is best given in parentheses immediately after the heading. No column should contain the same number or unit throughout its length. Footnotes to tables are denoted by superscript a, b, c... The units used should be clearly stated. Confusion can arise from the use of powers in column headings. The following usage is recommended: e.g., if molar absorptivities are listed, the heading should be  $\epsilon(10^4 \text{ l mol}^{-1} \text{ cm}^{-1})$  so that a number 2.32 in the column signifies 23 200.

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Basic SI and other accepted metric nomenclature are given in the Appendix. In accordance with IUPAC rules, the mass number, atomic number, number of atoms and ionic charge should be designated by a left upper index, a left lower index, a right lower index and a right upper index, respectively, placed round the atomic symbol. For example, the phosphate ion should be designated as  $\text{PO}_4^{3-}$  (not  $\text{PO}_4^{-3}$  or  $\text{PO}_4^{---}$ ), and phosphorus-32 as  $^{32}\text{P}$  (not  $\text{P}^{32}$  or  $\text{P-32}$ ).

The Stock notation for the indication of stoichiometric valency states (and indirectly the proportion of the constituents) should be used. Examples are iron(III) chloride rather than ferric chloride, and potassium hexacyanoferrate(II) rather than potassium ferrocyanide. These rules are valid for French and German as well as English usage.

The use of nanometre (nm) and micrometre ( $\mu\text{m}$ ) for the expression of analytical wavelengths has long superseded  $m\mu$  or  $\text{\AA}$  or  $\mu$ , all of which should be avoided, although  $\text{\AA}$  is sensibly retained in crystallographic work.

Natural or Napierian logarithms should be denoted by  $\ln$  and decadic logarithms by  $\log$ .

Molarity ( $\text{mol l}^{-1}$  or  $M$ ) is the preferred concentration unit, but normality ( $N$ ) can be used for convenience if it does not introduce ambiguity.

Unusual abbreviations require definition when first used. Abbreviations for long chemical names (e.g., EDTA, HEDTA, TBAH, en, pn, Tris) are useful, especially in equations, tables or figures. For ease of distinction, well-known techniques may be abbreviated by using lower-case letters and full stops, such as, g.c.-m.s., u.v., i.r., a.a.s.,  $^{13}\text{C}$ -n.m.r., a.s.v., d.p.p., etc. In the interests of clarity, however, excessive use of abbreviations is not encouraged.

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Decimal points should be indicated by full stops in papers written in English and by commas in French and German papers. All decimal numbers smaller than unity should include a leading zero (e.g., 0.11).

## Appendix

### Basic SI units

metre	m	candela	cd
kilogram	kg	mole	mol
second	s	(an Avogadro number of particles such as atoms, molecules, ions, electrons.)	
ampere	A		
degree Kelvin	K		

### Derived SI units

joule	J	$\text{kg m}^2 \text{s}^{-2}$	farad	F	$\text{A s V}^{-1}$
newton	N	$\text{J m}^{-1}$	weber	Wb	$\text{V s}$
watt	W	$\text{J s}^{-1}$	henry	H	$\text{V s A}^{-1}$
coulomb	C	$\text{A s}$	tesla	T	$\text{V s m}^{-2}$
volt	V	$\text{J A}^{-1} \text{s}^{-1}$	hertz	Hz	$\text{s}^{-1}$
ohm	$\Omega$	$\text{V A}^{-1}$	degree Celsius	$^{\circ}\text{C}$	$\text{K} - 273.15$

### Other units

litre	l	$10^{-3} \text{ m}^3$	hour	h	$3.6 \times 10^3 \text{ s}$
gram	g	$10^{-3} \text{ kg}$	dyne	dyn	$10^{-5} \text{ N}$
poise	P	$10^{-3} \text{ m}^{-1} \text{ s}^{-1}$	atmosphere	atm	$101.325 \text{ kN m}^{-2}$
electron volt	eV	$1.6021 \times 10^{-19} \text{ J}$	molar	M	$\text{mol l}^{-1}$
calorie	cal	$4.184 \text{ J}$	molal	m	$\text{mol kg}^{-1}$
minute	min	60 s	curie	Ci	$3.7 \times 10^{10} \text{ s}^{-1}$

Prefixes to abbreviations for the names of units indicating

Multiples		Sub-multiples			
tera ( $\times 10^{12}$ )	T	milli ( $\times 10^{-3}$ )	m	pico ( $\times 10^{-12}$ )	p
giga ( $\times 10^9$ )	G	micro ( $\times 10^{-6}$ )	$\mu$	femto ( $\times 10^{-15}$ )	f
mega ( $\times 10^6$ )	M	nano ( $\times 10^{-9}$ )	n	atto ( $\times 10^{-18}$ )	a
kilo ( $\times 10^3$ )	k				



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