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- 1 W. Lund and M. Salberg, *Anal. Chim. Acta*, 76 (1975) 131.
- 2 M. McDaniel, A. D. Shendrikar, K. D. Reizneir and P. W. West, *Anal. Chem.*, 48 (1976) 2240.

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Books

- 1 D. D. Perrin, *Masking and Demasking of Chemical Reactions*, Interscience-Wiley, New York, 1970, p. 188.
- 2 S. Hofmann, in G. Svehla (Ed.), *Wilson and Wilson's Comprehensive Analytical Chemistry*, Vol. 9, Elsevier, Amsterdam, 1979, p. 89.

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Extraction (%)	95.0	99.8	99.5	89.0

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The Stock notation for the indication of stoichiometric valency states (and indirectly the proportion of the constituents) should be used. Examples are iron(III) chloride rather than ferric chloride, and potassium hexacyanoferrate(II) rather than potassium ferrocyanide. These rules are valid for French and German as well as English usage.

The use of nanometre (nm) and micrometre (μm) for the expression of analytical wavelengths has long superseded $m\mu$ or \AA or μ , all of which should be avoided, although \AA is sensibly retained in crystallographic work.

Natural or Napierian logarithms should be denoted by \ln and decadic logarithms by \log .

Molarity (mol l^{-1} or M) is the preferred concentration unit, but normality (N) can be used for convenience if it does not introduce ambiguity.

Unusual abbreviations require definition when first used. Abbreviations for long chemical names (e.g., EDTA, HEDTA, TBAH, en, pn, Tris) are useful, especially in equations, tables or figures. For ease of distinction, well known techniques may be abbreviated by using lower-case letters and full stops, such as, g.c.-m.s., u.v., i.r., a.a.s., ^{13}C -n.m.r., a.s.v., d.p.p., etc. In the interests of clarity, however, excessive use of abbreviations is not encouraged.

Ambiguity in expressing dilution can be avoided by the use of, e.g., (1 + 2) rather than 1:2 which could mean either one part diluted with two parts or one part diluted to twice its volume.

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Decimal points should be indicated by full stops in papers written in English and by commas in French and German papers. All decimal numbers smaller than unity should include a leading zero (e.g., 0.11).

Appendix

Basic SI units

metre	m	candela	cd
kilogram	kg	mole	mol
second	s	(an Avogadro number of	
ampere	A	particles such as atoms,	
degree Kelvin	K	molecules, ions, electrons.)	

Derived SI units

joule	J	$\text{kg m}^2 \text{s}^{-2}$	farad	F	A s V^{-1}
newton	N	J m^{-1}	weber	Wb	V s
watt	W	J s^{-1}	henry	H	V s A^{-1}
coulomb	C	A s	tesla	T	V s m^{-2}
volt	V	$\text{J A}^{-1} \text{s}^{-1}$	hertz	Hz	s^{-1}
ohm	Ω	V A^{-1}	degree Celsius	$^{\circ}\text{C}$	$\text{K} - 273.15$

Other units

litre	l	10^{-3}m^3	hour	h	$3.6 \times 10^3 \text{s}$
gram	g	10^{-3}kg	dyne	dyn	10^{-5}N
poise	P	$10^{-3} \text{m}^{-1} \text{s}^{-1}$	atmosphere	atm	101.325kN m^{-2}
electron volt	eV	$1.6021 \times 10^{-19} \text{J}$	molar	M	mol l^{-1}
calorie	cal	4.184 J	molar	m	mol kg^{-1}
minute	min	60 s	curie	Ci	$3.7 \times 10^{10} \text{s}^{-1}$

Prefixes to abbreviations for the names of units indicating

Multiples		Sub-multiples			
tera ($\times 10^{12}$)	T	milli ($\times 10^{-3}$)	m	pico ($\times 10^{-12}$)	p
giga ($\times 10^9$)	G	micro ($\times 10^{-6}$)	μ	femto ($\times 10^{-15}$)	f
mega ($\times 10^6$)	M	nano ($\times 10^{-9}$)	n	atto ($\times 10^{-18}$)	a
kilo ($\times 10^3$)	k				

INFORMATION FOR AUTHORS

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Journals

- 1 W. Lund and M. Salberg, *Anal. Chim. Acta*, 76 (1975) 131.
 - 2 M. McDaniel, A.D. Shendrikar, K.D. Reizneir and P. W. West, *Anal. Chem.*, 48 (1976) 2240.
- The title of the journal must be abbreviated as in the Bibliographic Guide for Editors and Authors.

Books

- 1 D.D. Perrin, *Masking and Demasking of Chemical Reactions*, Interscience-Wiley, New York, 1970, p. 188.
- 2 S. Hofmann, in G. Svehla (Ed.), *Wilson and Wilson's Comprehensive Analytical Chemistry*, Vol. 9, Elsevier, Amsterdam, 1979, p. 89.

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